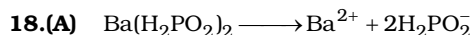
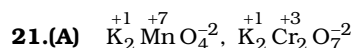
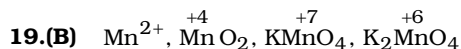


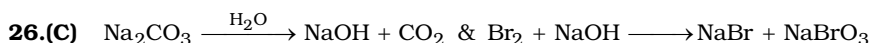
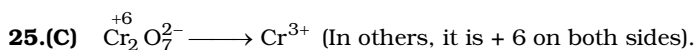
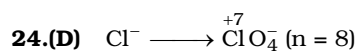
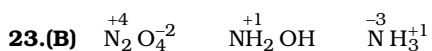
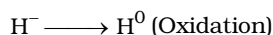
17.(BD) NaNO_2 and Na_2O_2 because one of the element is in intermediate oxidation state.



For $\text{H}_2\text{PO}_2^- \Rightarrow +2 + x - 4 = -1 \Rightarrow x = +1$



22.(D) Increase is O.N $\equiv$ oxidation and Decrease is O.N $\equiv$ Reduction.



On balancing and simplifying the equation : $3\text{Br}_2 \xrightarrow{+6\text{OH}^-} 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$



$\therefore \text{g meq of MnO}_4^- = \text{g meq of M}^{x+}$

$1 \times 5 = 2.5 \times (5 - x) \Rightarrow x = 3$

